

Original Research Article

**Theoretical insights of the electronic structures, conductivity,
and aromaticity of graphyne and Si-doped graphynes**

Reza Ghiasia^{a,*}, Faegheh Aghazadeh Kozeh Kanani^b

^a Department of Chemistry, East Tehran Branch, Islamic Azad University, Tehran, Iran

^b Department of Chemistry, Science and Research Branch, Islamic Azad University, Tehran, Iran

ARTICLE INFORMATION

Received: 28 August 2022
Received in revised: 25 September 2022
Accepted: 10 October 2022
Available online: 01 December 2022

DOI: 10.26655/AJNANOMAT.2022.4.4

KEYWORDS

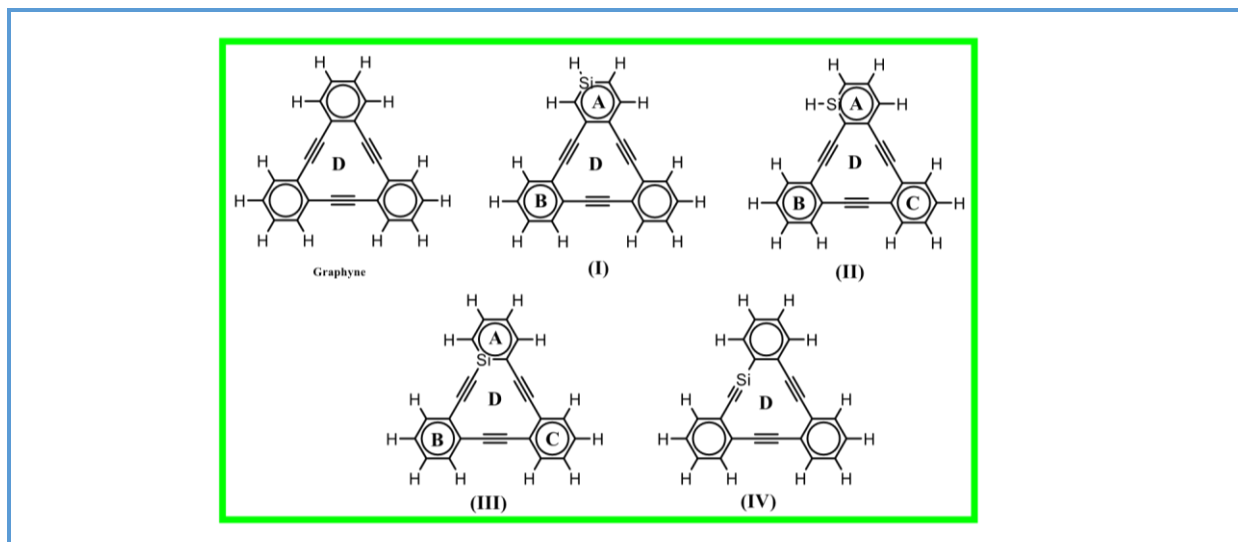
Graphyne
Si-Doped Graphyne
Nucleus Independent Chemical Shift (NICS)
Electron Localization Function (ELF)

ABSTRACT

In the present research, the dipole moment, electronic structure, frontier orbitals energy, and aromaticity in the graphyne and Si-doped graphynes were studied with M06-2X quantum chemical computation. The relative energies of four possible isomers of Si-doped graphyne were calculated. Also, the ionization potential (IP) and electron affinity (EA) values of the studied molecules were reported. HOMO-LUMO gap values were used for illustration of these molecules conductivity. Aromaticity of these molecules was investigated by nucleus independent chemical shift (NICS) values and electron localization function (ELF).

© 2022 by SPC (Sami Publishing Company), Asian Journal of Nanoscience and Materials, Reproduction is permitted for noncommercial purposes.

Graphical Abstract



Introduction

Graphyne which is a newkind of artificial carbon allotropes is correlated to graphite/graphene. The special structures and attractive optical, electronic and mechanical properties have been explored. Graphyne was first depicted in 1987 [1], while, in spite of numerous attempts, graphyne has not yet been prepared [2-5]. Theoretical researches excited attempts for synthesis of graphyne layers, crystals, and nanostructures. Graphyne compounds were synthesized by the polymerization of molecules having carbon cage fragments that are similar to the corresponding structure type of graphyne [6, 7]. In this regard, theoretical calculations of the structures of graphyne layer fragments and single graphyne layers and their properties were reported [8-15]. The major aim of the present investigation is to clarify the Sidoping on the structure and properties of graphyne by the quantum mechanics method. The dipole moment, structural parameters, frontier orbital energies and aromaticity of these molecules were explored.

Computational Methods

All calculations were carried out using Gaussian 09 suite program [16]. The calculations of systems contain C, Si and H described by the standard 6-311G(d,p) basis set. Geometry optimization was performed utilized by the hybrid meta exchange correlation functional of Truhlar and Zhao (M06-2X) [17]. A vibrational analysis was performed at each stationary point which confirmed its identity as an energy minimum.

After optimization of the studied molecules, single point calculations were carried out at the MP2/6-311G(d,p) [18-22] and wB97XD/6-311G(d,p) [23, 24] levels of theory. Chemical shift values are calculated using Gauge independent atomic orbital (GIAO) method at the same method and basis sets of optimization [25]. The topological analysis of the ELF was made using the Multiwfn 3.3.5 [26] at the M062X/6-311G(d,p) level of theory.

Results and Discussion

Energetic aspects

Figure 1 presents the structures and numbering schemes of atoms in graphyne and isomers of Si-doped graphynes. The absolute energies and relative energies of these molecules are listed in Table 1. As shown in Table 1, the III-isomer is the most stable isomer of Si-doped graphynes.

IV isomer in the planar form has one imaginary frequency which exhibits a bending vibration of out-plane for Si atom. Therefore, the stable form of this isomer deviates from planarity, changes to nonplanar form and also changes the C_s symmetry of isomer to C_1 .

Table 1. Energy (E, in a.u) and relative energy (ΔE , kcal/mol) of graphyne and Si-doped graphyne isomers at the determined levels of theory.

	E ^a	ΔE^a	E ^b	ΔE^b	E ^c	ΔE^c
Graphyne						
Si-doped Graphyne	-921.4726	-				
I-Isomer	- 1172.8112	12.03	-1170.057978	9.815127	- 1172.8441763	12.18931
II-Isomer	- 1172.8133	10.70	- 1170.0600007	8.545864	- 1172.8461199	10.96968
III-Isomer	- 1172.8303	0.00	- 1170.0736194	0	- 1172.8636012	0
IV-Isomer	- 1172.7807	31.15	- 1170.0240762	31.08883	- 1172.8100839	33.58261
I-Isomer	- 1172.8112	12.03	-1170.057978	9.815127	- 1172.8441763	12.18931
II-Isomer	- 1172.8133	10.70	- 1170.0600007	8.545864	- 1172.8461199	10.96968

^aoptimization at the M062X/6-311G(d,p) level of theory.

^bafter optimization at the M062X/6-311G(d,p) level of theory, single point of optimized structure of the calculated at MP2/6-311G(d,p) level of theory ^c after optimization at the M062X/6-311G(d,p) level of theory, single point of optimized structure of the calculated at wb97xd/6-311G(d,p) level of theory

Moreover, the results of single point optimized studied molecules show that the calculations at the MP2/6-311G(d,p) and III-

isomer is the most stable isomer of Siwb97xd/6-311G(d,p) levels of theory of doped graphynes.

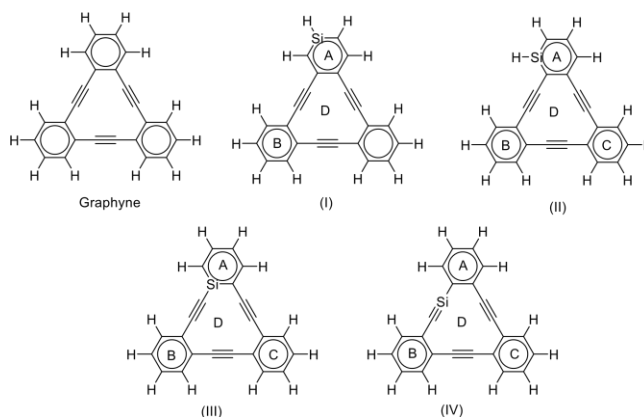


Figure 1. The structures of graphyne and Si-doped graphyne isomers in this study

Polarizability

Polarizability which describes the response of a system in an applied electric field determines not only the strength of molecular interactions (such as the long range intermolecular induction, dispersion forces, etc.) and the cross sections of different scattering and collision processes, but also the nonlinear optical properties of the system [27].

The isotropic polarizability (α_{iso}) is calculated as the mean value as given in the following equation:

$$\alpha_{iso} = \langle \alpha \rangle = \frac{(\alpha_{xx} + \alpha_{yy} + \alpha_{zz})}{3}$$

and the polarizability anisotropy invariant (α_{aniso}) is:

$$\alpha_{aniso} = \left[\frac{(\alpha_{xx} - \alpha_{yy})^2 + (\alpha_{yy} - \alpha_{zz})^2 + (\alpha_{zz} - \alpha_{xx})^2}{2} \right]^{\frac{1}{2}}$$

The isotropic and anisotropic polarizability values of studied molecules are calculated (Table 1). These values show the smallest isotropic and anisotropic polarizability values for **III**-isomer. As expected from the principles of minimum energy (MEP), and minimum polarizability (MMP), that is, while a conformer changes from the most stable to other less stable species in most cases, the energy

Table 2. Dipole moment (μ , Debye), isotropic polarizability (α_{iso} , Bohr³), anisotropic polarizability (α_{aniso} , Bohr³), frontier orbital energies (a.u) and HOMO-LUMO gap (eV) of graphyne and Si-doped graphyne isomers in the M062X/6 311G(d,p) level of theory.

	μ	α_{iso}	α_{aniso}	E(HOMO)	E(LUMO)	Gap
Graphyne	0.00	291.04	299.59	-0.2518	-0.0537	5.39
Si-doped Graphyne						
I-Isomer	1.22	320.38	326.04	-0.2405	-0.0548	5.05
II-Isomer	0.67	318.31	321.93	-0.2473	-0.0564	5.19
III-Isomer	2.12	315.50	316.97	-0.2386	-0.0605	4.85
IV-Isomer	1.24	324.27	276.08	-0.2479	-0.0847	4.44

increases, and the polarizability increases[28-30].

IPs and EAs

The ionization potential (IP) and electron affinity (EA) are the well-defined properties that can be calculated by DFT to estimate electrons into the compounds. These values aim to get an in-detail rationalization of the relationship between the structure and the electronic behavior of the molecule, in particular the response of the molecule to the formation of a hole, or to the addition of an electron, additional information is derived. Table 2 contains the calculated IPs, EAs, both vertical and adiabatic, and the extraction potentials (HEP and EEP for the hole and electron, respectively) that refer to the geometry of the ions [31, 32].

The results, presented in Table 2, show that the IP_a and IP_v values in Si-doped gaphynes are less than graphyne. The most stable isomer of Si-doped gaphynes (**III**-isomer) has the lowest IP_a and IP_v values. On the other hand, the EA_a and EA_v values in Sidoped gaphynes are more than graphyne. The most stable isomer of Si-doped gaphynes (**III**-isomer) has the most EA_a and EA_v values.

The mobility of charges has been established to be related mainly to the internal reorganization energy $\lambda_{\text{hole/electron}}$. The data in Table 2 show that the λ_{hole} values for all molecules are all smaller than their corresponding $\lambda_{\text{electron}}$ values indicating that the electron transfer rate is lower than the hole transfer rate. According to the values gathered in Table 2, these molecules can be applied as hole transport layer (HTL)/an electron block layer.

Dipole moments

The dipole moments of ingraphyne and isomers of Si-doped graphynes are gathered in Table 1. As can be seen, the most stable isomer of Si-doped graphynes (III-isomer) has the largest dipole moment than that of other isomers.

Molecular orbital analysis

Table 3. Ionization potential (IP, in eV), electron affinity (EA, in eV), hole Extraction potential (HEP, eV), electron extraction potential (EEP, eV), reorganization energies (\square_{hole} , $\square_{\text{electron}}$, eV) of graphyne and Si-doped graphyne isomers in the M062X/6-311G(d,p) level of theory.

	IP(a)	EA(a)	IP(v)	EA(v)	HEP	EEP	\square_{hole}	$\square_{\text{electron}}$
Graphyne	7.366	0.990	7.472	0.844	7.255	1.139	0.217	0.295
Si-doped Graphyne								
I-Isomer	6.993	1.022	7.117	0.894	6.865	1.156	0.252	0.262
II-Isomer	7.156	1.089	7.298	0.943	7.009	1.242	0.289	0.299
III-Isomer	6.949	1.196	7.080	1.046	6.816	1.351	0.264	0.305

The HOMO-LUMO gap is regarded as a (HOMO) and the conduction level (LUMO) major determinant of the electrical conductivity of a system. It is defined as the calculated using the following equation energy difference between the valance level [37]:

$$\sigma \propto \exp\left(-\frac{E_g}{2k_b T}\right)$$

Table 1 summarizes the values of frontier orbital energies and HOMO-LUMO gaps. Apparently, HOMO is more stable in graphyne than Si-doped graphynes. In contrast, LUMO is more stable in Si-doped graphynes than graphyne.

The hardness (η), chemical potential (μ) and electrophilicity (ω) of the molecules are evaluated by the following formulas [33-36]:

$$\eta = \frac{IP_a - EA_a}{2}$$

$$\mu = \frac{EA_a + IP_a}{2}$$

$$\omega = \frac{\mu^2}{2\eta}$$

Table 3 shows the largest hardness and smallest electrophilicity for II-isomer.

Where E_g , σ , k_b and T are HOMO-LUMO gap, electrical conductivity, Boltzmann's constant and T is the temperature, respectively. The HOMO-LUMO gap decreases from 2.69 eV in graphyne to 2.522.22 eV after Si-doping. According to the above mentioned equation, at a given temperature, the smaller HOMO-LUMO gap values are associated with more electrical conductivity.

Figure 2 presents the plots of frontier orbitals of the studied molecules. It can be seen that these orbitals are π -type, and the overlap

between carbon atoms of $\text{-C}\equiv\text{C-}$ is bonding and anti-bonding for HOMO and LUMO, respectively.

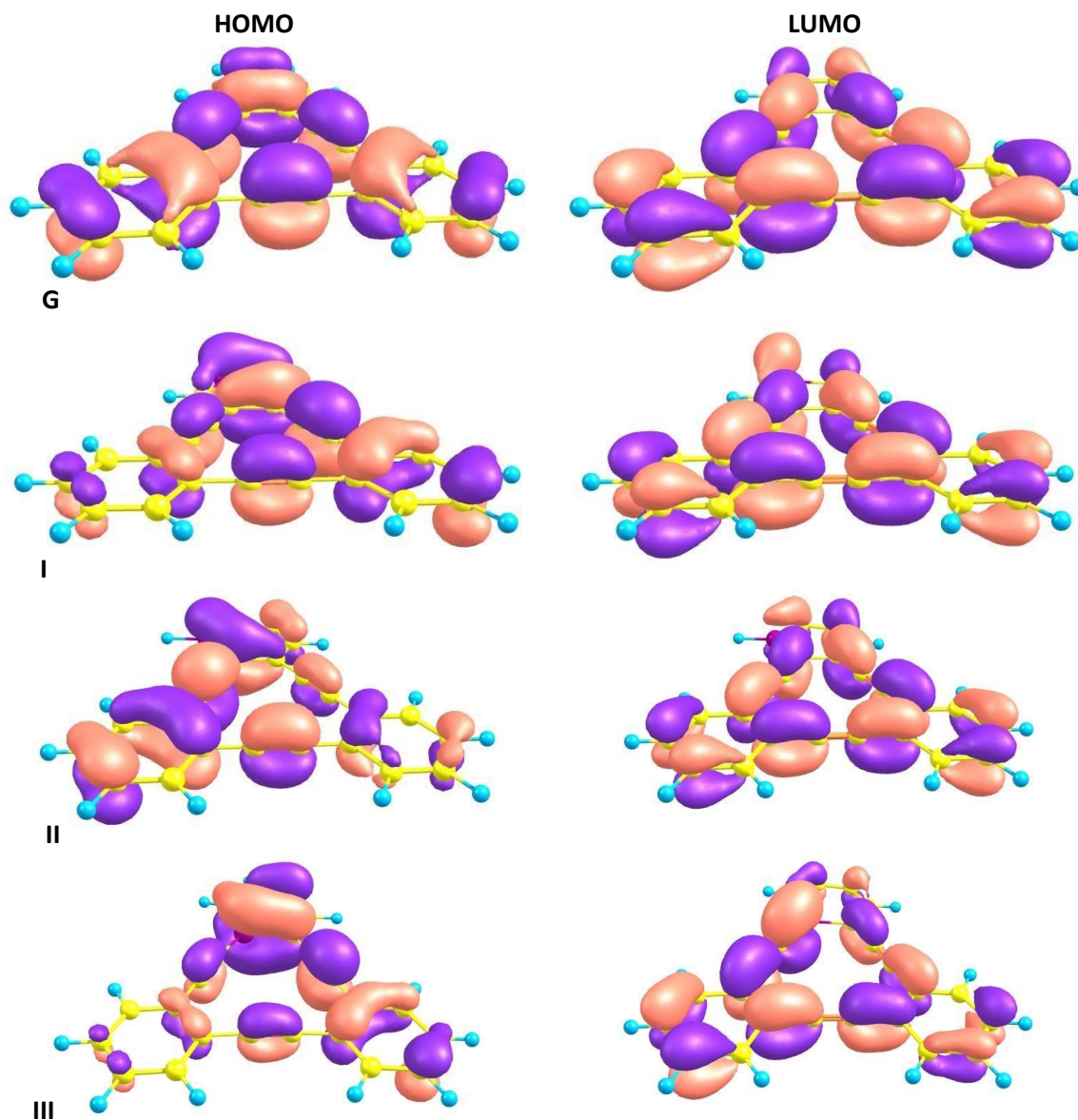


Figure 2. The plots of frontier orbitals in graphyne and Si-doped graphyne isomers

Chemical shift nucleus independent values (NICS)

Figure 1 illustrates that graphyne involves aromatic 6-membered rings and the weak anti-aromatic 12-membered rings[38, 39]consisting of alternating triple bonds and conjugated double bonds, the latter being shared by the 6-

membered rings. As benzene is regarded to be the smallest unit of graphite, dehydrobenzo[12]annulene (1,DBA)[40] can be considered as the smallest unit of graphyne. It is, therefore, of interest to illustrate changes of aromaticity in Si-doped graphynes.

In an aromatic ring, a diatropic ring current rises which is the result of magnetic shielding at the ring center, however in an anti-aromatic ring, a paratropic ring current outcomes in deshielding at the ring center [41, 42]. Chemical shift nucleus independent values (NICS) values give a suitable and valuable quantity of the degree of aromaticity/anti-aromaticity of one ring in a molecule [43].

Chemical shift nucleus independent values (NICS) values are calculated for the study of

Table 4. Hardness (η , eV), chemical potential (μ , eV) and electrophilicity (ω , eV) values of of graphyne and Si-doped graphyne isomers in the M062X/6-311G(d,p) level of theory

	η	μ	ω
Graphyne	3.19	4.18	2.74
Si-doped Graphyne			
I-Isomer	2.99	4.01	2.69
II-Isomer	3.03	4.12	2.80
III-Isomer	2.88	4.07	2.88

Electron localization function (ELF)

An alternative scheme based on properties of the electron density, as the electron localization function (ELF) [44], has been introduced to understand aromaticity [45, 46]. The electron

aromaticity in the investigated molecules. These values are gathered in Table 4. The calculated NICS values show the negative values for **A**, **B**, **C** cycles. These values are compatible with aromaticity of these rings. The most negative values in 1.5 Å which are above the rings reveal π -aromaticity in the rings. But, the positive values are observed for **D** cycle. Therefore, there is an anti-aromaticity in the **D** cycle.

delocalization can be obtained through the bifurcation analysis of ELF. The bifurcation points have been interpreted as a measure of and, chemically, as a measure of electron the interaction among the different basins delocalization [45].

Table 5. NICS (in ppm) and ELF bifurcation values of rings for graphyne and Si-doped graphyne isomers in the M062X/6 311G(d,p) level of theory.

		0.0	0.5	1.5	2.0	ELF
Graphyne	A	-6.76	-8.79	-9.52	-7.08	0.949
	D	3.07	3.01	2.35	1.73	-
Si-doped Graphyne						
I-Isomer	A	-7.12	-7.99	-8.01	-6.12	0.700
	B	-6.58	-8.63	-9.46	-7.09	0.932
	C	-6.90	-8.93	-9.57	-7.06	0.941
	D	2.70	2.70	2.23	1.66	-
II-Isomer	A	-6.57	-7.58	-7.73	-5.97	0.701
	B	-6.58	-8.60	-9.39	-7.01	0.924
	C	-7.12	-9.16	-9.69	-7.09	0.941
	D	2.63	2.66	2.27	1.70	-
III-Isomer	A	-6.24	-7.17	-7.44	-5.86	0.702
	B	-6.65	-8.70	-9.48	-7.07	0.943
	C	-6.55	-8.57	-9.35	-6.95	0.941
	D	3.35	3.30	2.57	1.93	-

The ELF values in the ring critical point (RCP) of six-membered rings are reported in Table 4. It can be found, high bifurcation values of ELF in the studied molecules. The ELF bifurcation values of the Si-doped graphyne are lower than that of garphyne molecule. This fact can be attributed to the electronegativity differences between carbon and silicon atoms.

Conclusion

We used M06-2X method and 6-311G(d,p) basis set in order to investigate the structure and properties of the graphyne and Si-doped graphyne. These investigations indicated:

1. The comparison of energy values for four isomers of Si-doped graphyne shows that **III**-isomer is the most stable isomer.
2. The internal reorganization energy values illustrate that these molecules can be applied as hole transport layer (HTL)/an electron block layer.
3. The HOMO-LUMO gap values of the Si-doped graphynes are less than graphyne, therefore it can be seen as having more electrical conductivity in Si-doped graphynes.
4. The NICS values of these studied molecules explore π -aromaticity in the A, B, C rings. In contrast, antiaromaticity in the **D** cycle can be found.
5. Smaller ELF bifurcation values of the Si-doped graphyne as compared to garphyne molecule are attributed to the electronegativity differences between carbon and silicon atoms.

Disclosure Statement

No potential conflict of interest was reported by the authors.

References

- [1]. Baughman R.H., Eckhardt H., Kertesz M., *J. Chem. Phys.*, 1987, **87**:6687 [[Crossref](#)], [[Google Scholar](#)], [[Publisher](#)]
- [2]. Yoshimura T., Inaba A., Sonoda M., Tahara K., Tobe Y., Williams R.V., *Org. Lett.*, 2006, **8**:2933 [[Crossref](#)], [[Google Scholar](#)], [[Publisher](#)]
- [3]. Li G.X., Li Y.L., Liu H.B., Guo Y.B., Li Y.J., Z. D.B., *Chem. Commun.*, 2010, **46**:3256 [[Crossref](#)], [[Google Scholar](#)], [[Publisher](#)]
- [4]. Qian X., Ning Z., Li Y., Liu H., Ouyang C., Chen Q., Li Y., *Dalton Trans*, 2012, **41**:730 [[Crossref](#)], [[Google Scholar](#)], [[Publisher](#)]
- [5]. Wang S., Yi L.X., Halpert J.E., Lai X.Y., Liu Y.Y., Cao H.B., Yu R.B., Wang D., Li Y.L. *Small*, 2012, **8**:265 [[Crossref](#)], [[Google Scholar](#)], [[Publisher](#)]
- [6]. Diederich F., Rubin Y., *Angew. Chem. Int. Ed. Engl.*, 1992, **31**:1101 [[Crossref](#)], [[Google Scholar](#)], [[Publisher](#)]
- [7]. Diederich F., *Nature (London)*, 1994, **369**:199 [[Crossref](#)], [[Google Scholar](#)], [[Publisher](#)]
- [8]. Belenkov E.A., Chelyab I., *Nauchn. Tsentra UrO Ross. Akad. Nauk*, 2002, **14**:12.
- [9]. Coluci V.R., Braga S.F., Legoas S.B., Galvao D.S., Baughman R.H., *Phys. Rev. B: Condens. Matter*, 2003, **68**:35430 [[Crossref](#)], [[Google Scholar](#)], [[Publisher](#)]
- [10]. A.N. Enyashin, A.A. Sofronov, Y.N. Makurin, A.L. Ivanovski, *J. Mol. Struct.: Theochem.*, 2004, **684**:29 [[Crossref](#)], [[Google Scholar](#)], [[Publisher](#)]
- [11]. Coluci V.R., Galvao D.S., Baughman R.H. *J. Chem. Phys.*, 2004, **121**:3228 [[Crossref](#)], [[Google Scholar](#)], [[Publisher](#)]
- [12]. Coluci V.R., Braga S.F., Legoas S.B., Galvao D.S., Baughman R.H., *Nanotechnology*, 2004, **15**:S142 [[Crossref](#)], [[Google Scholar](#)], [[Publisher](#)]
- [13]. Lepetit C., Zou C., Chauvin R. *J. Org. Chem.*, 2006, **71**:6317 [[Crossref](#)], [[Google Scholar](#)], [[Publisher](#)]

- [14]. Belenkov E.A., Shakhova I.V. *Phys. Solid State*, 2011, **53**:2385 [[Crossref](#)], [[Google Scholar](#)], [[Publisher](#)]
- [15]. Podlivaev A.I., Openov L.A. *Phys. Solid State* 2012, **54**:1723 [[Crossref](#)], [[Google Scholar](#)], [[Publisher](#)]
- [16]. Frisch M.J., Trucks G.W., Schlegel H.B., Scuseria G.E., Robb M.A., Cheeseman J.R., Scalman G., Barone V., Mennucci B., Petersson G.A., Nakatsuji H., Caricato M., Li X., Hratchian H.P., Izmaylov A.F., Bloino J., Zheng G., Sonnenberg J.L., Hada M., Ehara M., Toyota K., Fukuda R., Hasegawa J., Ishida M., Nakajima T., Honda Y., Kitao O., Nakai H., Vreven T., Montgomery J.A., Peralta J.E.Jr., Ogliaro F., Bearpark M., Heyd J.J., Brothers E., Kudin K.N., Staroverov V.N., Kobayashi R., Normand J., Raghavachari K., Rendell A., Burant J.C., Iyengar S.S., Tomasi J., Cossi M., Rega N., Millam J.M., Klene M., Knox J.E., Cross J.B., Bakken V., Adamo C., Jaramillo J., Gomperts R., Stratmann R.E., Yazyev O., Austin A.J., Cammi R., Pomelli C., Ochterski J.W., Martin R.L., Morokuma K., Zakrzewski V.G., Voth G.A., Salvador P., Dannenberg J.J., Dapprich S., Daniels A.D., Farkas O., Foresman J.B., Ortiz J.V., Cioslowski J., Fox D.J. in, Gaussian, Inc., Wallingford CT, 2009 [[Google Scholar](#)]
- [17]. Zhao Y., Truhla D.G. *J. Phys. Chem. A*, 2006, **110**:5121 [[Crossref](#)], [[Google Scholar](#)], [[Publisher](#)]
- [18]. Head-Gordon M., Pople J.A., Frisch M.J. *Chem. Phys. Lett.*, 1988, **153**:503 [[Crossref](#)], [[Google Scholar](#)], [[Publisher](#)]
- [19]. Saebø S., Almlöf J. *Chem. Phys. Lett.*, 1989, **154**:83 [[Crossref](#)], [[Google Scholar](#)], [[Publisher](#)]
- [20]. Frisch M.J., Head-Gordon M., Pople J.A. *Chem. Phys. Lett.*, 1990, **166**:275 [[Crossref](#)], [[Google Scholar](#)], [[Publisher](#)]
- [21]. Frisch M.J., Head-Gordon M., Pople J.A. *Chem. Phys. Lett.*, 1990, **166**:281 [[Crossref](#)], [[Google Scholar](#)], [[Publisher](#)]
- [22]. Head-Gordon M., Head-Gordon T. *Chem. Phys. Lett.*, 1994, **220**:122 [[Crossref](#)], [[Google Scholar](#)], [[Publisher](#)]
- [23]. Grimme S. *J. Comput. Chem.*, 2006, **27**:1787 [[Crossref](#)], [[Google Scholar](#)], [[Publisher](#)]
- [24]. Chai J.D., Head-Gordon M. *Phys. Chem. Chem. Phys.*, 2008, **10**:6615 [[Crossref](#)], [[Google Scholar](#)], [[Publisher](#)]
- [25]. Wolinski K., Hinton J.F., Pulay P. *J. Am. Chem. Soc.*, 1990, **112**:8251 [[Crossref](#)], [[Google Scholar](#)], [[Publisher](#)]
- [26]. Lu T., Chen F. *J. Mol. Graphics. Model*, 2012, **38**:314 [[Crossref](#)], [[Google Scholar](#)], [[Publisher](#)]
- [27]. Librando V., Alparone A., Minniti Z. *J. Mol. Struct.*, 2008, **856**:105 [[Crossref](#)], [[Google Scholar](#)], [[Publisher](#)]
- [28]. Chattaraj P.K., Fuentealba P., Gomez B., Contreras R. *J. Am. Chem. Soc.*, 2000, **122**:348 [[Crossref](#)], [[Google Scholar](#)], [[Publisher](#)]
- [29]. Chattaraj P.K., Sengupta S., *J. Phys. Chem. A*, 1996, **100**:16126 [[Crossref](#)], [[Google Scholar](#)], [[Publisher](#)]
- [30]. Ghanty T.K., Ghosh S.K., *J. Phys. Chem. A*, 1996, **100**:12295 [[Crossref](#)], [[Google Scholar](#)], [[Publisher](#)]
- [31]. Curioni A., Boero M., Andreoni W. *Chem Phys Lett* 1998, **294**:263 [[Crossref](#)], [[Google Scholar](#)], [[Publisher](#)]
- [32]. Wang I., Botzung-Appert E., Stéphan O., Ibanez A., Baldeck P.L. *J. Opt. A*, **2002**, **4**:S258 [[Crossref](#)], [[Google Scholar](#)], [[Publisher](#)]
- [33]. Pearson R.G., *Chemical Hardness*; Wiley-VCH: Oxford, 1997 [[Crossref](#)], [[Google Scholar](#)], [[Publisher](#)]
- [34]. Parr R.G., Yang W., *DensityFunctional Theory of Atoms and Molecules*; Oxford University Press: New York, 1989 [[Crossref](#)], [[Google Scholar](#)], [[Publisher](#)]
- [35]. Parr R.G., Szentpály L.v., Liu S. *J. Am. Chem. Soc.*, 1999, **121**:1922 [[Crossref](#)], [[Google Scholar](#)], [[Publisher](#)]

- [36]. Pearson R.G. *Proc. Natl. Acad. Sci.*, 1986, **83**:8440 [[Crossref](#)], [[Google Scholar](#)], [[Publisher](#)]
- [37]. Li S. *Semiconductor Physical Electronic*; 2nd ed., Springer: USA, 2006 [[Google Scholar](#)], [[Publisher](#)]
- [38]. Wiberg K.B. *Chem. Rev.*, 2001, **101**, 1317 [[Crossref](#)], [[Google Scholar](#)], [[Publisher](#)]
- [39]. Spitler E.L., Johnson C.A., Haley M.M. *Chem. Rev.*, 2006, **106**:5344 [[Crossref](#)], [[Google Scholar](#)], [[Publisher](#)]
- [40]. Staab H.A., Graf F. *Tetrahedron Lett.*, 1966, **7**:751 [[Crossref](#)], [[Google Scholar](#)], [[Publisher](#)]
- [41]. Pauling L., *J. Chem. Phys.*, 1936, **4**:673 [[Crossref](#)], [[Google Scholar](#)], [[Publisher](#)]
- [42]. Pople J.A., Untch K.G., *J. Am. Chem. Soc.*, 1966, **88**:4811 [[Crossref](#)], [[Google Scholar](#)], [[Publisher](#)]
- [43]. Chen Z., Wannere C.S., Corminboeuf C., Puchta R., Schleyer P.V.R. *Chem. Rev.*, 2005, **105**:3842 [[Crossref](#)], [[Google Scholar](#)], [[Publisher](#)]
- [44]. Becke A., Edgecombe K., *J. Chem. Phys.*, 1990, **92**:5397 [[Crossref](#)], [[Google Scholar](#)], [[Publisher](#)]
- [45]. Savin A., Becke A., Flad D., Nesper R., Preuss H., Schnering H.V., *Angew. Chem., Int. Ed. Engl.*, 1991, **30**:409 [[Crossref](#)], [[Google Scholar](#)], [[Publisher](#)]
- [46]. Santos J.C., Tiznado W., Contreras R., Fuentealba P. *J. Chem. Phys.*, 2004, **120**:1670 [[Crossref](#)], [[Google Scholar](#)], [[Publisher](#)]

How to cite this manuscript: Reza Ghiasi*, Faegheh Aghazadeh Kozeh Kanani. Theoretical insights of the electronic structures, conductivity, and aromaticity of graphyne and Si-doped graphynes. *Asian Journal of Nanoscience and Materials*, 2022, 5(4), 303-312. DOI: 10.26655/AJNANOMAT.2022.4.4